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Constant Expression, Initial

Concentration, K_p , K_c ,

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constant can then be determined for each mixture; the average should be the equilibrium constant value for the formation of the FeSCN^{2+} ion. In Part A of this experiment, you will prepare FeSCN^{2+} solutions

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System. In this video you
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The equilibrium constant, K , is used to quantify the equilibrium state. The expression for the equilibrium constant for a reaction is determined by

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Examining the balanced
chemical equation.

~~Experiment 3 Determination
of an Equilibrium Constant
for ...~~

The experiment was completed
to determine the equilibrium

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Lab Answers constant of a chemical reaction This chemical reaction created a state of chemical equilibrium. than one, or small, less than one, but not negative or zero. A large K_c (equilibrium constant) means

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~~Lab Answers~~ that at equilibrium, the concentrations of the products will generally be greater than the

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About The Equilibrium
Constant~~

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: Calculate the

concentration of SCN⁻ at equilibrium for Trials 1-4

using the equation: $[SCN^-]$

$eq = [SCN^-]_{i} - [FeSCN_2^+]_{eq}$

7. Calculate K_c for Trials

1-4. Be sure to show the K_c expression and the values

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Lab Answers substituted in for each of these calculations. 8. Using your four calculated K_c values, determine an average value for K_c . How constant were your K_c

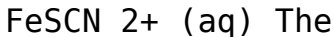
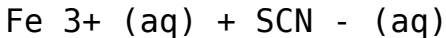
~~Lab 1: Chemical Equilibrium:~~

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~~Finding a Constant, Kc~~



The equilibrium constant for equation 4 can be expressed using the concentrations of the three components: (5)

$$K = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}]}$$

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[SCN -] In order to calculate the equilibrium constant, one must simultaneously determine the concentrations of all three of the components.

~~Lab 5 - Determination of an~~

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Question: Pre-Lab – The
Determination Of An

Equilibrium Constant DATA

TABLE Parts I And II Beaker

[FeSCN₂⁺] Absorbance 1 0.00

M 0.001 2 8.0 X 10⁻⁵ M 0.238

3 1.2 X 10⁻⁴ M 0.344 4 1.6 X

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10⁻⁴ M 0.435 5 2.0 X 10⁻⁴ M
0.530 Unknown, Part II 0.302
Insert Excel Absorbance Vs
Concentration Curve Here
Include Linear Least Squares
Fit Too 1. (Part II) Use The
Calibration ...

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7/21/17 Experiment 34 Post-
lab Questions 1. The
experiment I would conduct~~

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is to measure the absorbance of FeNCS_2^+ at various ranges, starting from 300nm to 500nm, to analyze UV-visible spectra. From that graph, if 447 nm is the optimal wavelength, the absorbance at that

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particular wavelength should be higher than the other. 2.

A) Since absorbance is proportional to ...

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100%. = $10^{-\epsilon bc} = 10^{-A}$.

Then, you will use your understanding of equilibrium processes to deduce the equilibrium concentrations of the reactants. Knowing all three concentrations listed above in Eq. 2. $K =$.

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[FeSCN²⁺] [Fe³⁺] [SCN⁻]
allows the equilibrium
constant for this reaction
to be calculated.

~~Lab 11 — Spectroscopic
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The lab handout is attached below in both Word and PDF format. (Note: Hat tip to Mr. Tony Locke, who shared this lab with an IB workshop I attended. I have modified it some along the way.) The

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Lab is a pretty basic look at causing disruptions to an equilibrium and making predictions - and then observations - based on Le Châtelier's Principle.

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