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$P_{\text{vap}} = \exp\left(\frac{14.790}{T} - \frac{5763}{T^2}\right)$ and by the ideal gas law
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 $n = \frac{P V}{R T} = \frac{18 \text{ mol} \times 8.314 \text{ J mol}^{-1} \text{ K}^{-1} \times 10^{-5} \text{ (bar} \cdot \text{m}^3 \text{ mol}^{-1} \text{ K}^{-1})}{T} = \frac{V}{V_1} \times 10^{-3} - 3$
 $- (10 - x) \times 10^6$ volume taken up by liquid
Equating P and P_{vap} we have $x = 18 \times 8.314 \times 10^{-5} T F$
 $5432.8 \times I$.

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Chapter 1 Solutions. Engineering and Chemical Thermodynamics. Wyatt Tenhaeff Milo Koretsky.

Department of Chemical Engineering Oregon State University. koretsm@enr.orst.edu 1.2 An

approximate solution can be found if we combine Equations 1.4 and 1.5: $1.2 \text{ mV} = \frac{2.3 \text{ kT}}{2}$

$2.3 \text{ kT} = V \cdot m$. Assume the temperature is 22 C .

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