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Aqueous Systems” The Water

Molecule: a Review. • Water is a simple tri-atomic molecule, H₂O.

• Each O-H bond is highly polar, because of the high electronegativity of the oxygen (N, O, F, and Cl have high values) • bond angle of water = 105°.

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- Page 501: 11 Answer All ionic
compounds are electrolytes because

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they dissociate into ions and conduct an electrical current in aqueous solution or in the molten state.

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colloidal particles, caused by collision with particles of the solvent in which they are dispersed. a compound that conducts an electric current when it is in an aqueous solution or in the molten state; all ionic compounds are electrolytes, but most covalent compounds are not.

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containing particles that are small enough to remain dispersed in the solvent and do not separate on standing.

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Chemistry, Chapter 15, Water and Aqueous Systems. surface tension. surfactant. aqueous solution. solvent. the inward force or pull that tends to minimize the surface ar.... any substance that interferes with hydrogen bonding between wa.... is

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water that contains dissolved
substances. the dissolving medium in
a solution.

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445–449) This section describes the properties of water in the liquid and solid states and explains how hydrogen bonding affects the surface tension and vapor pressure of water. Water in the Liquid State (pages 445–447) 1.

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SECTION 15.1 WATER AND ITS

PROPERTIES (pages 445–449)

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systems worksheet answers ,

source:eur-lex.europa.eu He may

want to stretch himself once a worker

knows his efforts do not go

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unnoticed. For instance, if he knows his performance will be judged based on achievement of a target, he will work harder to achieve it.

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crystal of solid NaCl is placed into an
aqueous NaCl solution. No precipitate
forms in the bottom of the container.
The final solution is: a.) super
saturated ... Chapter 5: Soil Water 57
Terms. hannah_lange5. Chapter 4:

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2. Section 15.1 Water and it ’ s

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Properties OBJECTIVES: – Explain the high surface tension and low vapor pressure of water in terms of the structure of the water molecule and hydrogen bonding.

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solvent and do not separate on
standing.

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Test B Answers

CHAPTER 15 | Aqueous Equilibria:
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Collect and Organize Figure

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P15 shows four lines to describe the possible dependence of percent ionization of acetic acid with concentration. We are to choose the one that best represents the trend for this weak acid. Analyze

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