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## Chapter 15 Review Acids Bases Answer Key

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-Lewis acid-base reaction is the formation of one or more covalent bonds between an electron-pair donor and an electron pair acceptor, although the 3 acid-base reactions differ, many compounds can be categorized as acids ...

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Chapter 15: Acids and Bases Chapter 15: Acids and Bases 1. The hydronium ion and the hydroxide ion, in that order, are: A)  $\text{H}_3\text{O}^+$ ,  $\text{OH}^-$  B)  $\text{OH}^-$ ,  $\text{H}_3\text{O}^+$  C)  $\text{OH}^-$ ,  $\text{H}^+$  D)  $\text{H}_3\text{O}^+$ ,  $\text{OH}^-$  E)  $\text{H}_3\text{O}^-$ ,  $\text{OH}^-$  Ans: D Category: Easy Section: 15.1 2. Which of the following does not fit the definition of a Brønsted Acid?

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Chapter 15 Acids and Bases. strong acid. strong base. weak acid. weak base. an acid that ionizes completely in solvent. a base that ionizes completely in a solvent. an acid that releases few hydrogen ions in aqueous solution. a base that releases few hydroxide ions in aqueous solution.

~~Chapter 15 Review Acids Bases 2 Answers~~

ACIDS AND BASES 453 SECTION 15-1 OBJECTIVES List five general properties of aqueous acids and bases. Name common binary acids and oxyacids, given their chemical formulas. List five acids commonly used in industry and the laboratory, and give two properties of each. Define acid and base according to Arrhenius' s theory of ionization. Explain the differences

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Chapter 15: Acids and Bases Acids and Bases Arrhenius

De fi nitions: acids - compounds that produce an increase in  $[H^+]$  when dissolved in water bases - compounds that produce an increase in  $[OH^-]$  when dissolved in water Lewis De fi nitions: acids - electron pair acceptors bases - electron pair donors Br ø nsted-Lowry De fi nitions: acids -  $H^+$  donors

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Thus,  $\text{H}_2\text{O}$  and  $\text{H}_3\text{O}^+$ , and  $\text{H}_2\text{O}$  and  $\text{OH}^-$  are conjugate acid-base pairs, but  $\text{H}_3\text{O}^+$  and  $\text{OH}^-$  are not conjugate acid-base pair. A Brønsted-Lowry acid-base reaction involves a competition between two bases for a proton, in which the stronger base ends up being the most protonated at equilibrium. In the reaction:  $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-(\text{aq})$ ,

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Chapter 15 Review Acids Bases Label them as acid, base, conjugate acid (ca) ad conjugate base (cb) Answer:  $\text{CH}_3\text{NH}_2$  (aq) is the base

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(an amine),  $\text{H}_2\text{O}(\text{l})$  acts as the acid here,  $\text{CH}_3\text{NH}_3^+(\text{aq})$  is the ca,  $\text{OH}^-(\text{aq})$  is the cb; Remember that  $\text{H}^+$  is just short hand for  $\text{H}_3\text{O}^+$  called hydronium ion.  $\text{H}^+$  doesn't really exist in water.

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Chapter 14 Review: Acids and Bases. Section 14.1 Bronsted Acids and Bases; Acids - molecules that can lose  $\text{H}^+$  (proton donors)

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making  $\text{H}_3\text{O}^+$  in water ... Since  $x$  is  $=[\text{H}^+]$  find the  $\text{pH} = -\log 7.0746 \times 10^{-3} = 2.15$  (we need 2 decimal places in the final answer)  
For more ...

## ~~Chapter 15 Review Acid Base Equilibria~~

CHAPTER 14 REVIEW Acids and Bases SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Name the following compounds as acids: sulfuric acid a.  $\text{H}_2\text{SO}_4$  sulfurous acid b.  $\text{H}_2\text{SO}_3$  hydrosulfuric acid c.  $\text{H}_2\text{S}$  perchloric acid d.  $\text{HClO}_4$  hydrocyanic acid e. hydrogen cyanide 2. H

## ~~14 Acids and Bases—David Brearley High School~~

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